Chemical Bonding Assessment

Naı	me:	Period:
1. 2.	Draw a diagram using the elements Na and Cl showing how they bond. Label the type bond of formed.	
3. 4.	Draw a diagram using the elements H and O showing how they bond. Label the type of bond formed.	
5.	Explain the Octet Rule.	

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6.	Electrons in the outermost orbital are ca	ılled	electrons.	
7.	An elements outermost orbital can hold	no more than	electrons.	
8.	Ionic bonds	electrons.		
9.	Covalent bonds	electrons.		
10.	Non-metal elements bond with metal elements	ements to create a/an		bond.
11.	Non-metal elements bond with non-met bond.	ral elements to create a/an		
12.	If an element has more electrons than pr	rotons, it will have a		_charge.
Lab	el the following as either ionic or molecu	ular compounds.		
13.	NaCl			
14.	O_2			
15.	CCl ₄			
16.	MgF_2	-		
Pre	dict the formular of the ionic compound	formed when the following	ng combine:	
17.	lithium and chlorine			
18.	calcium and oxygen			
19.	potassium and sulfur			
20.	potassium and sulfate (SO ₄ ²⁻)			

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